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~~Compounds | How to Pass Chemistry Gas Stoichiometry: Equations Part 1~~

~~Molar Conversions: Grams to Moles and Moles to Grams Limiting Reactant Practice Problem (Advanced) Converting Between Moles and Liters of a Gas at STP How to Use a Mole to Mole Ratio | How to Pass Chemistry Redox Reactions: Crash Course Chemistry #10 Mole Concept Stoichiometry - Limiting \u0026amp; Excess Reactant, Theoretical \u0026amp; Percent Yield - Chemistry Introduction to Combustion Analysis, Empirical Formula \u0026amp; Molecular Formula Problems~~

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GRAMS Molar mass (grams per mole) # MOLES 6.02 X 10²³ atoms or molecules per mole # MOLECULES 22.4 liters of gas per mole # LITERS (at STP) EXAMPLE 1: What would be the volume in liters of 40.36 grams of neon at STP? # liters Ne = 40.36 g Ne x 1 mole Ne x 22.4 L Ne = 44.80 L Ne 20.18 g Ne 1 mole Ne EXAMPLE 2: How many molecules would there be in 56 liters of carbon dioxide at STP? # molecules CO₂ = 56.0 L CO₂ x 1 mole CO₂ x 6.02 X 10²³ molecules CO₂ = 1.51 x 10²⁴ molecules CO₂ 22.4 L CO₂ 1 ...

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PAGE 3 CALCULATE THE MOLAR MASS FOR THE FOLLOWING COMPOUNDS OR DIATOMIC ELEMENTS. SET UP EACH PROBLEM AS SHOWN IN THE EXAMPLE ABOVE. INCLUDE UNITS (G/MOL) 1. water H_2O 2. calcium chloride $CaCl_2$ 3. copper(II) sulfate $CuSO_4$ 4. silver nitrate $AgNO_3$ 5. sulfuric acid H_2SO_4

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PAGE 6 CHEMISTRY WORKSHEET # 4 MOLAR MASSES

AND AVOGADRO'S NUMBER Now that you know two

definitions of a mole (a molar mass and an Avogadro's number of

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particles) you can combine these two definitions into one problem. EXAMPLE: How many molecules are there in 90.1 grams of water? $2 \text{ H} = 2 \times (1.01) = 2.02$ $\text{O} = 1 \times (16.00) = 16.00$ $2 \text{ H} + 1 \text{ O} = 18.02 \text{ g}$? molecules $\text{H}_2\text{O} = 90.1 \text{ g H}_2\text{O} \times \frac{1 \text{ mole H}_2\text{O}}{18.02 \text{ g}} \times 6.02 \times 10^{23} \text{ molecules H}_2\text{O} = 3.01 \times 10^{24} \text{ molecules H}_2\text{O}$...

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